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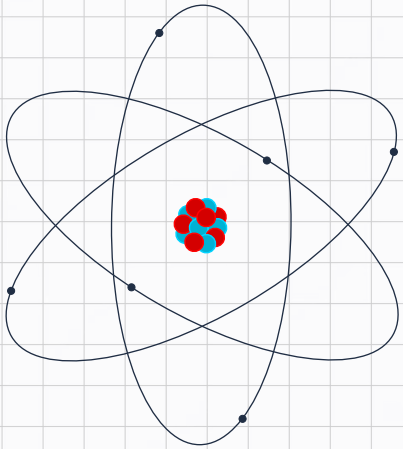
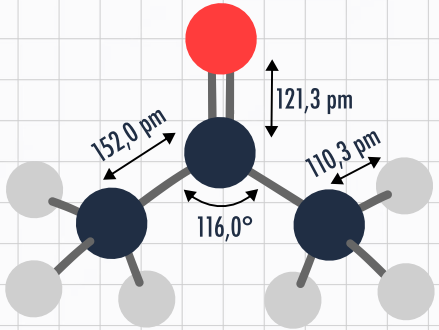


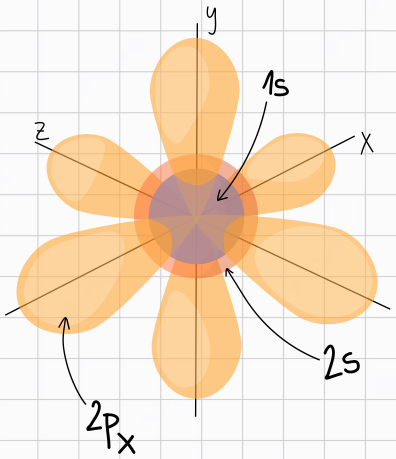
p^+ n^0 e^-

MOLEKULA



C_3H_6O

ORBITALY



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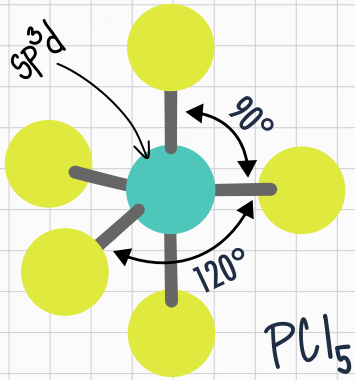

trigonální bipyramida

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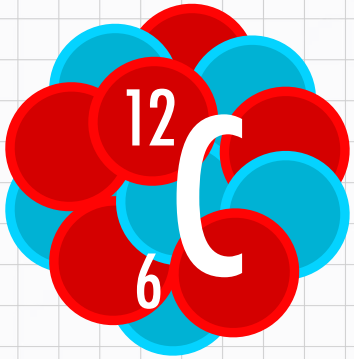

Z= 6 A=12

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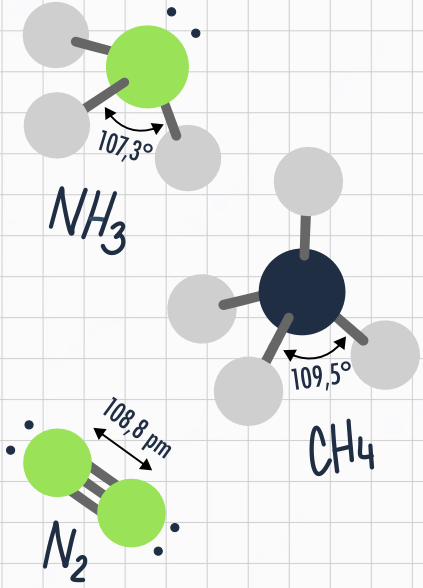

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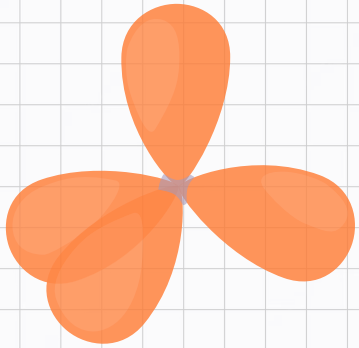
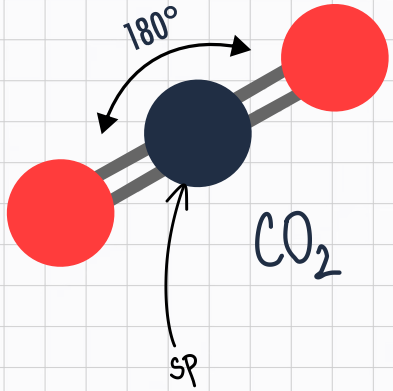


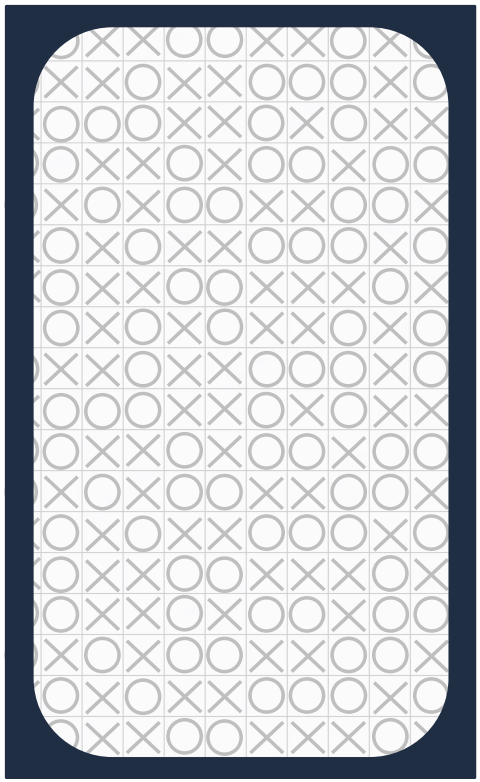
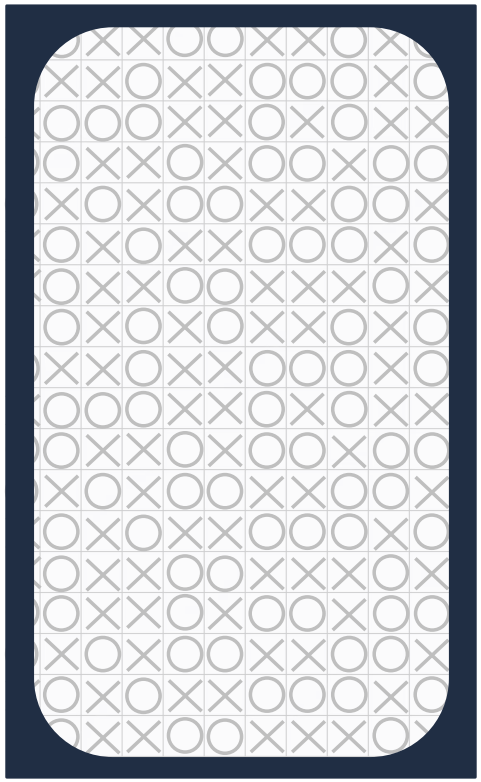
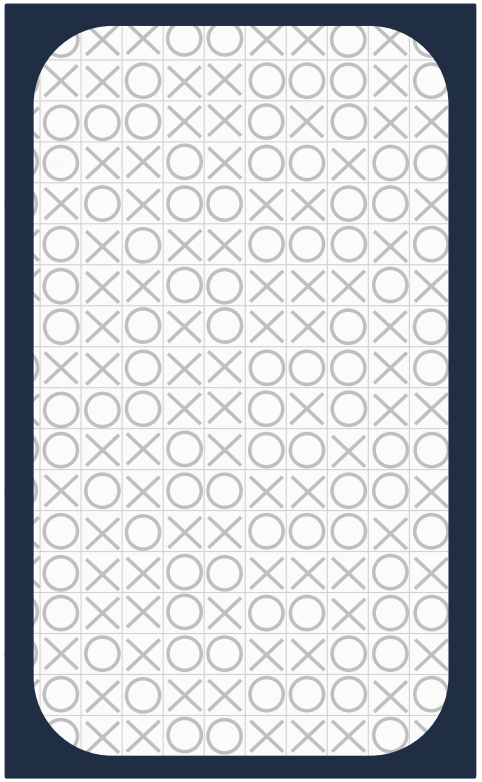
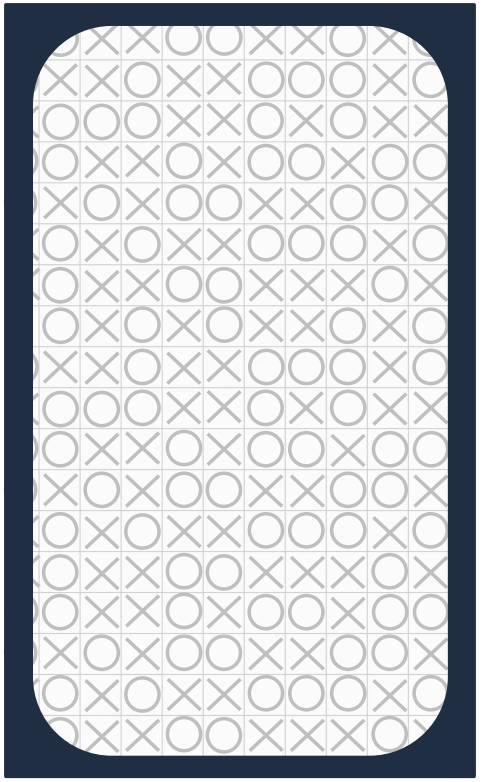
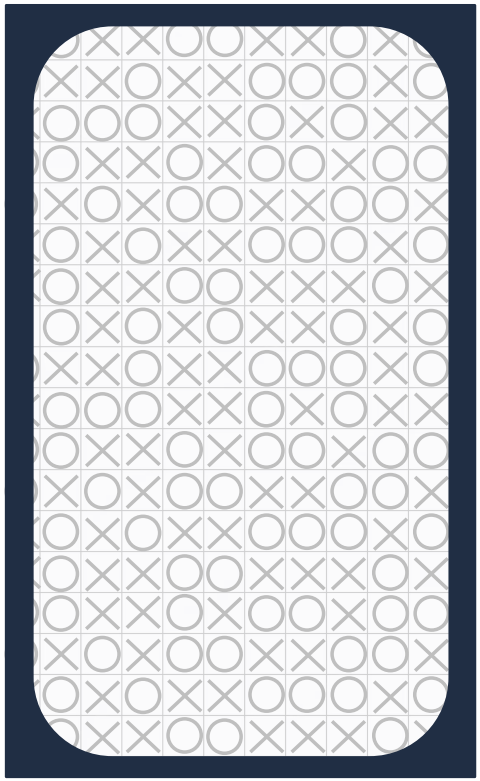
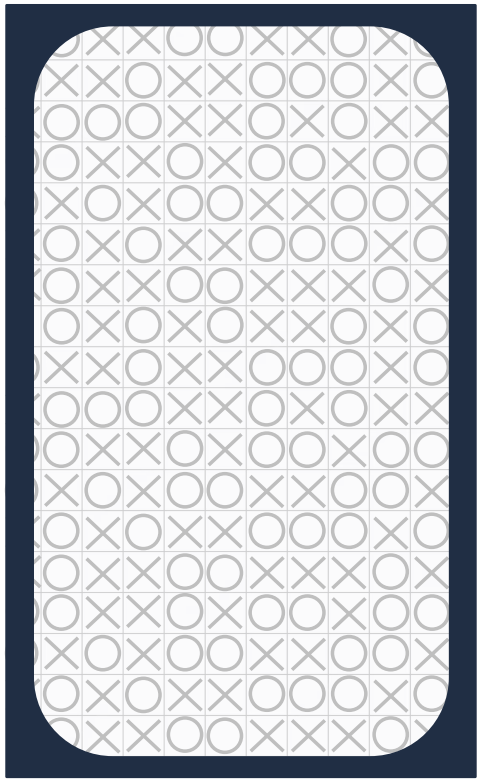
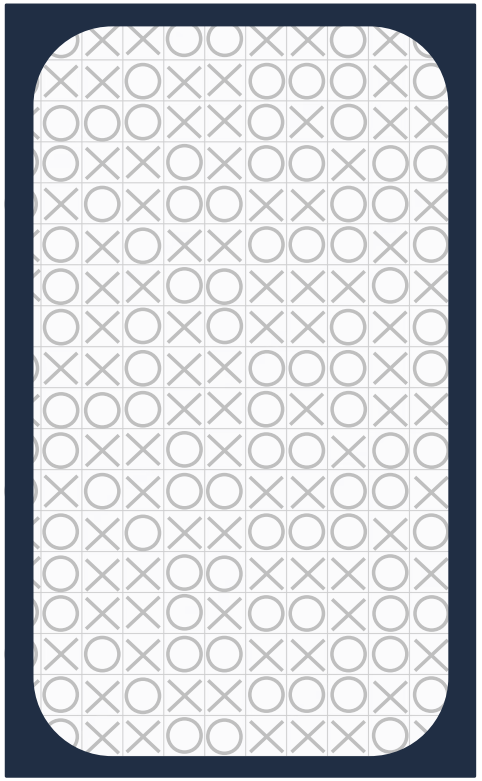
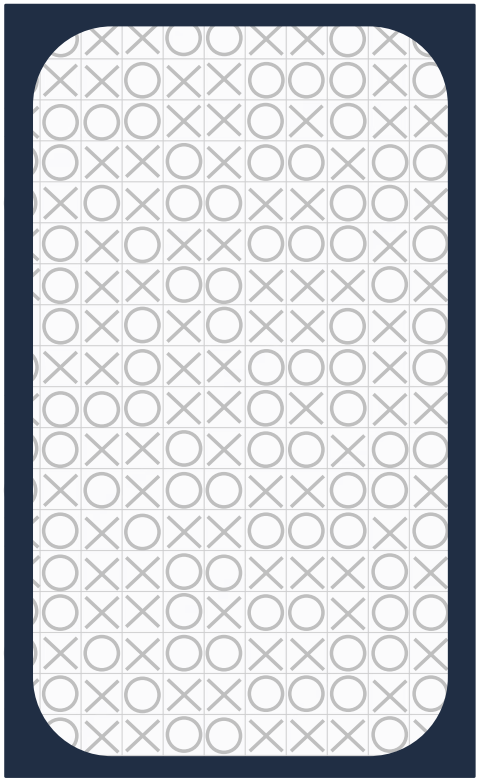
$2p_x + 2p_y + 2p_z + 2s = sp^3$

hybridizace orbitalů

VSEPR



lineární molekula

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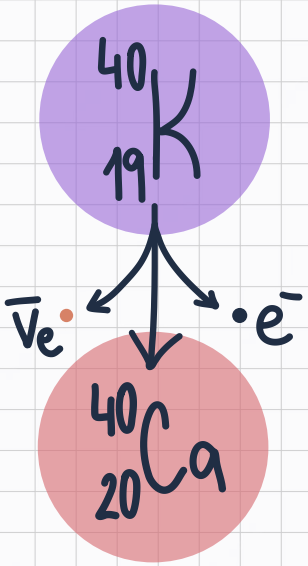
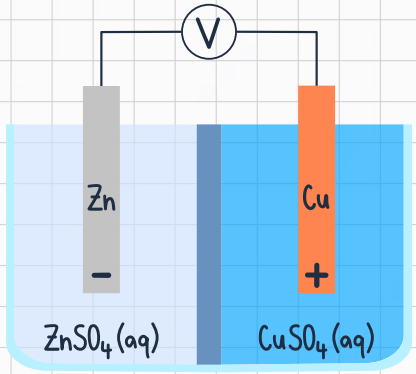
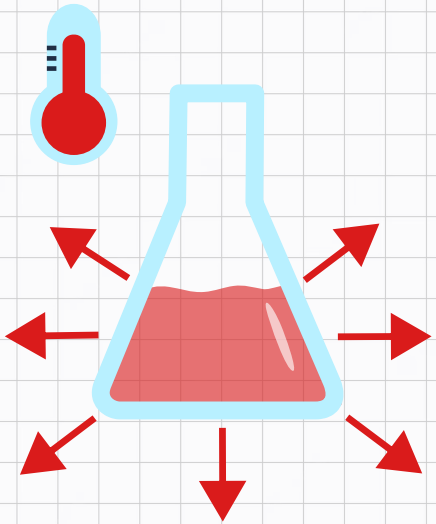


GALVANICKÝ ČLÁNEK



EXOTERMICKÁ REAKCE



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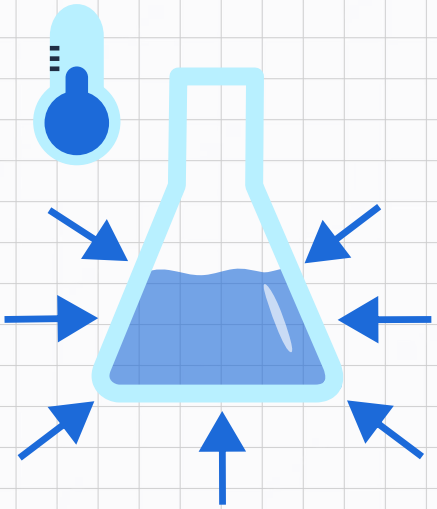

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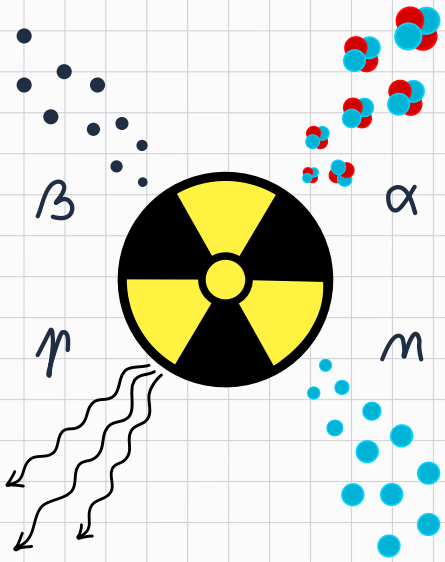

typy radioaktivního záření

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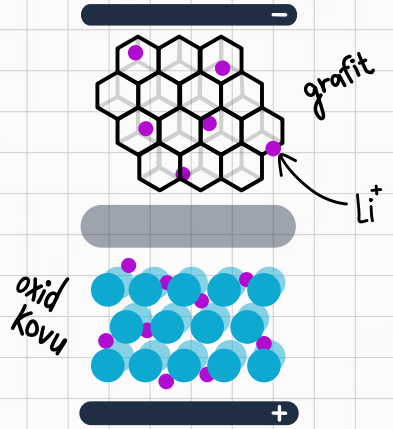
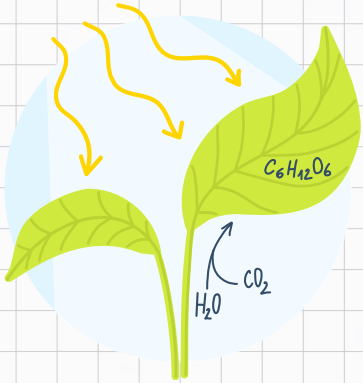


lithium-iontový akumulátor

EXOTERMICKÁ REAKCE

aluminotermie

ENDOTERMICKÁ REAKCE



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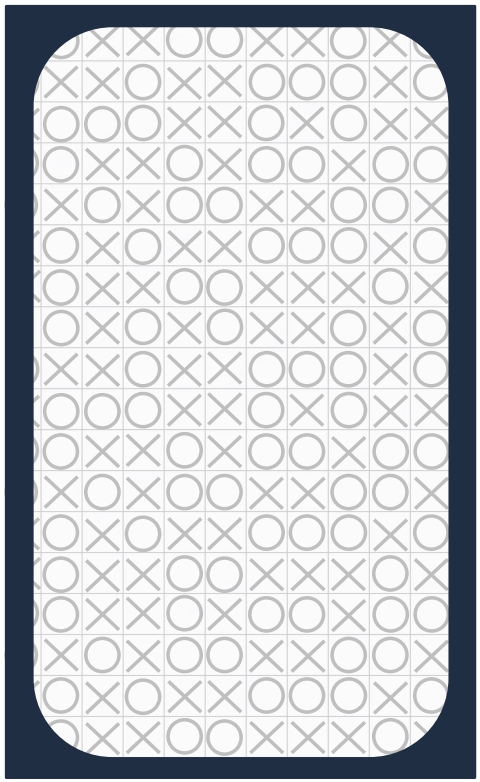
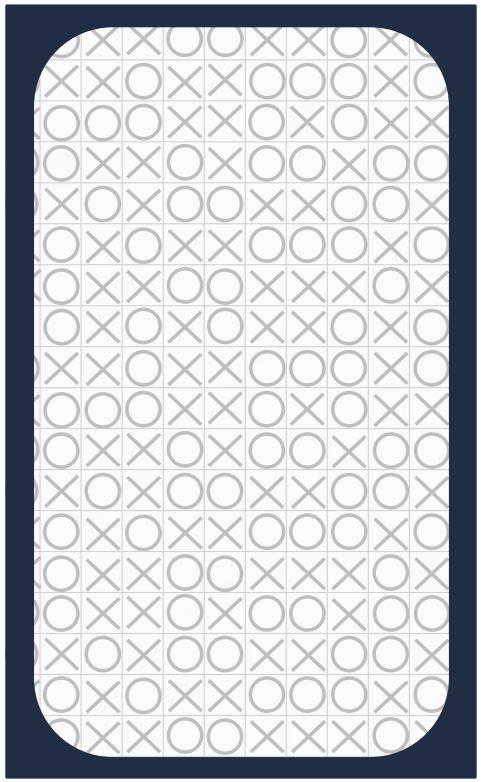
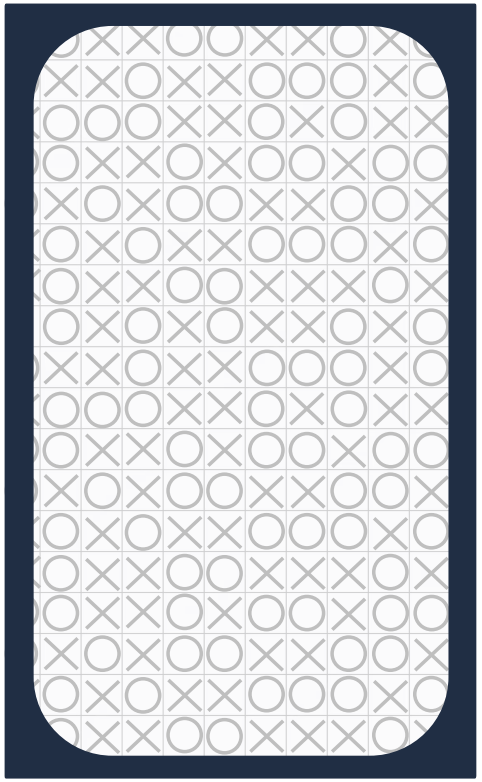
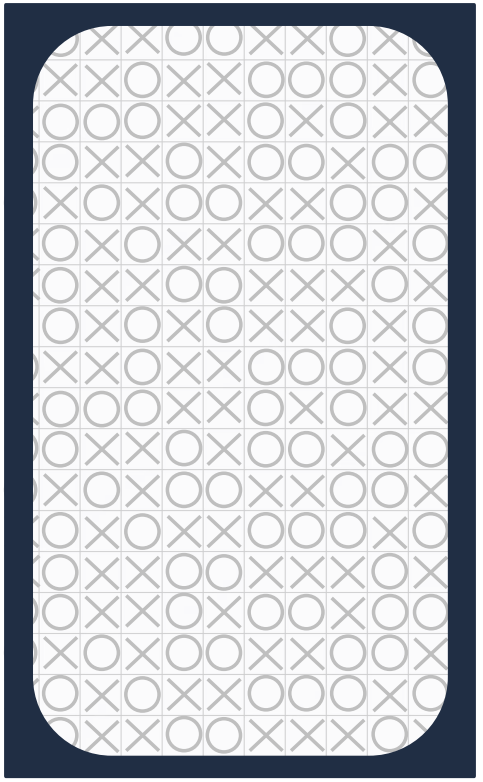
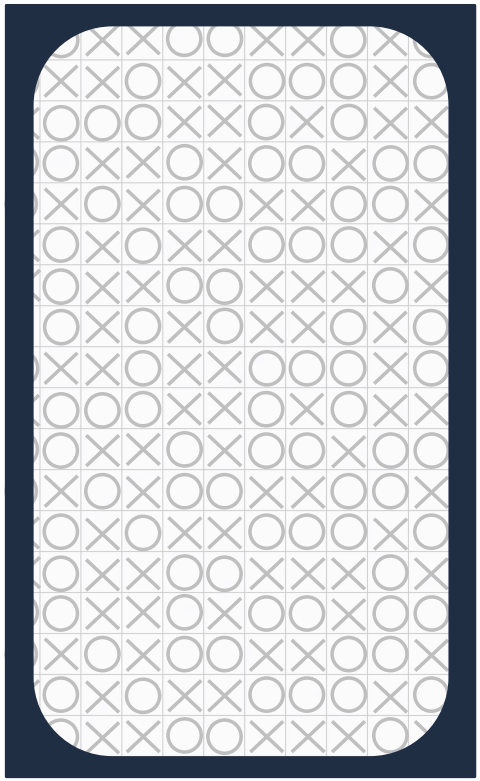
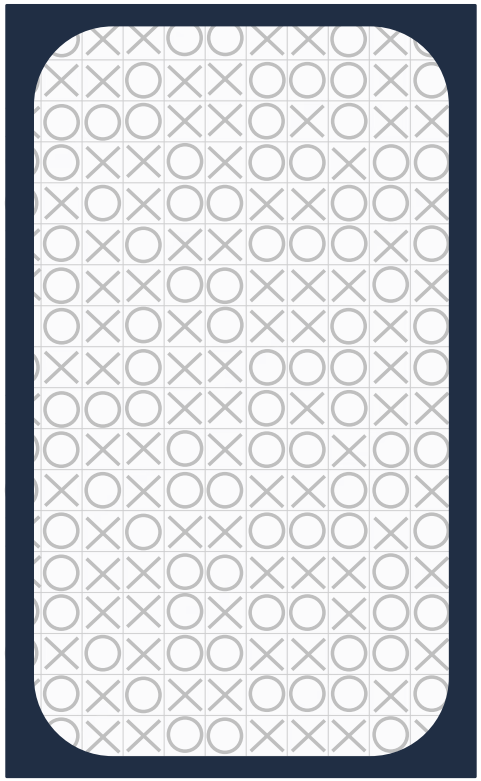
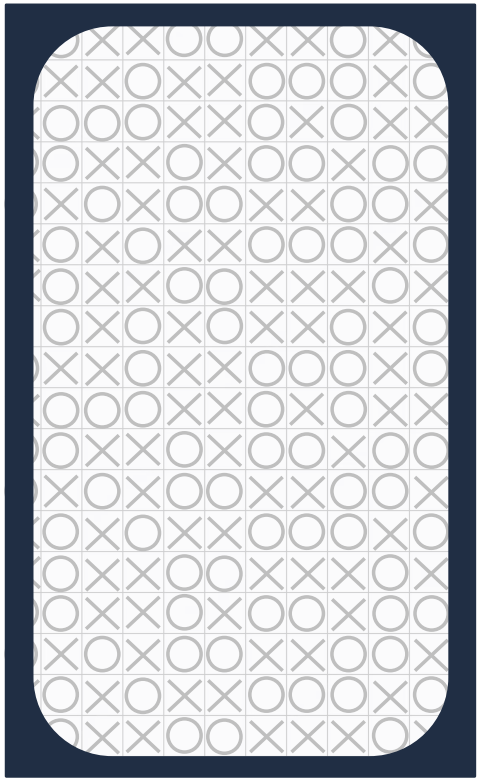
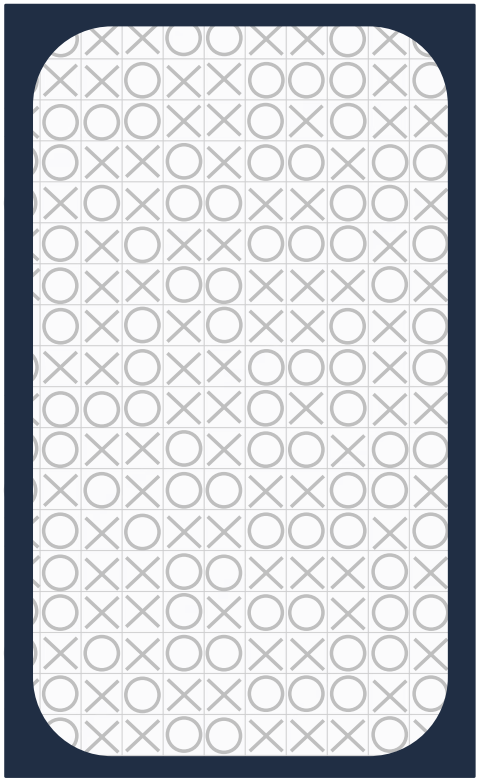
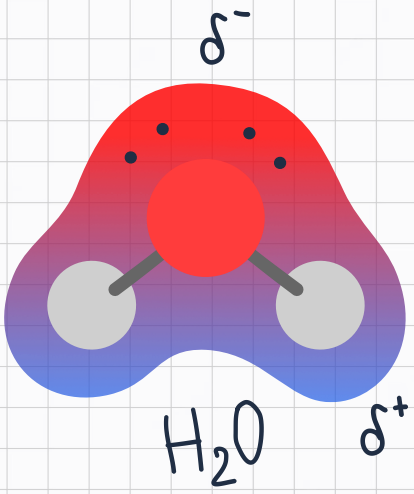









POLARITA VAZEB



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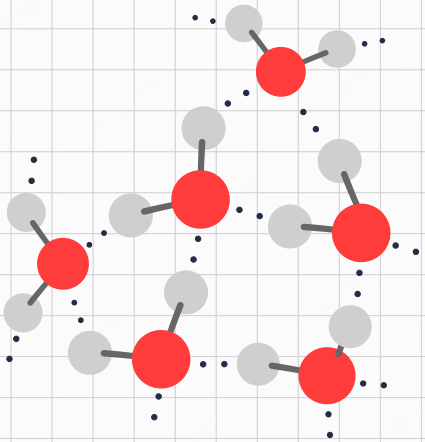
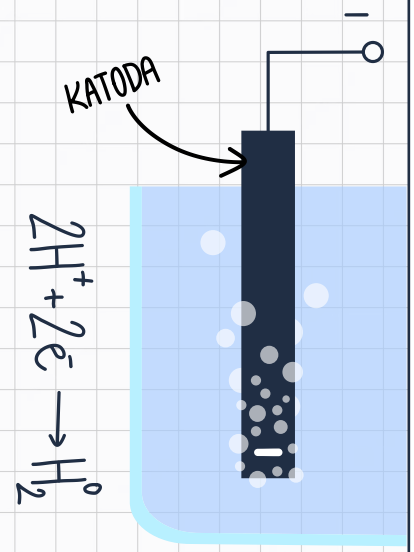


vodíkové můstky

REDUKCE



elektrolýza vody

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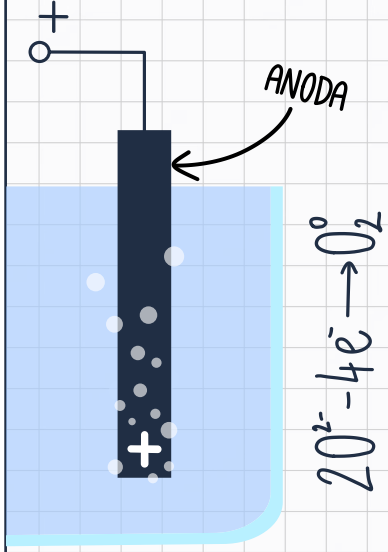

elektrolýza vody

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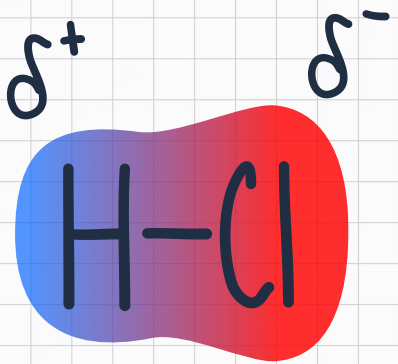
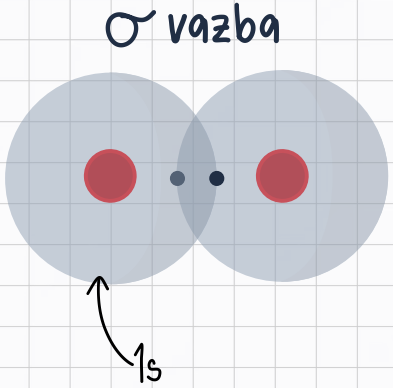


$X_H = 2{,}20$ $X_{Cl} = 3{,}16$

X elektronegativita

CHEMICKÁ VAZBA



překryv orbitalů

REDUKCE

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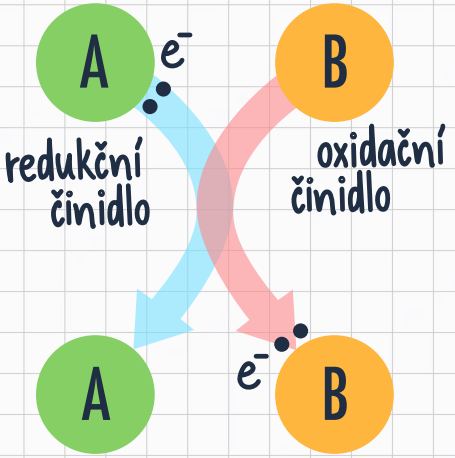
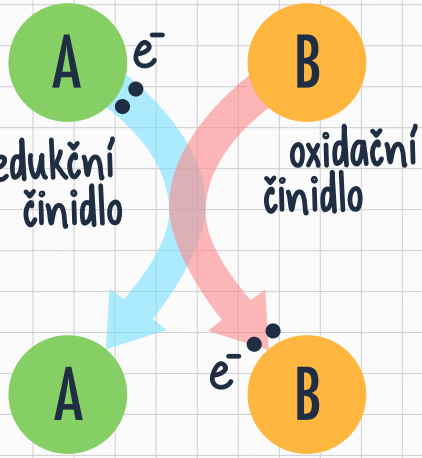
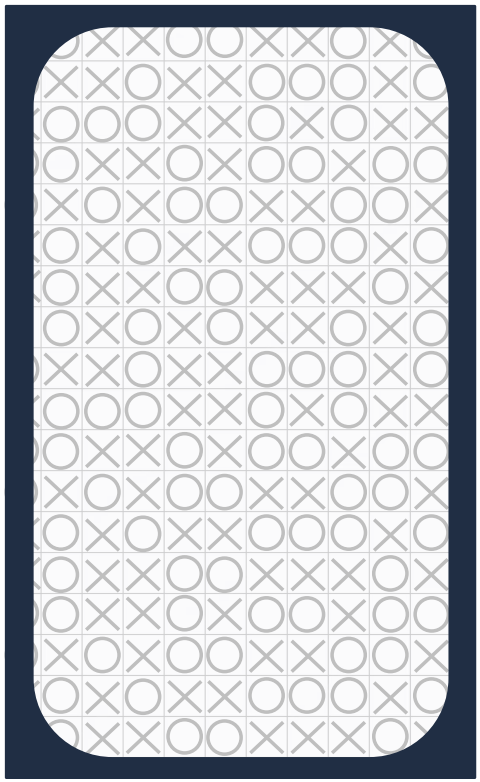
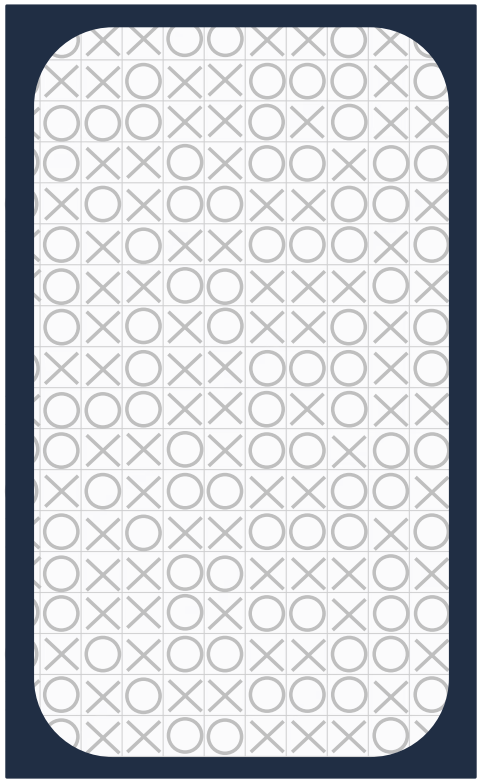
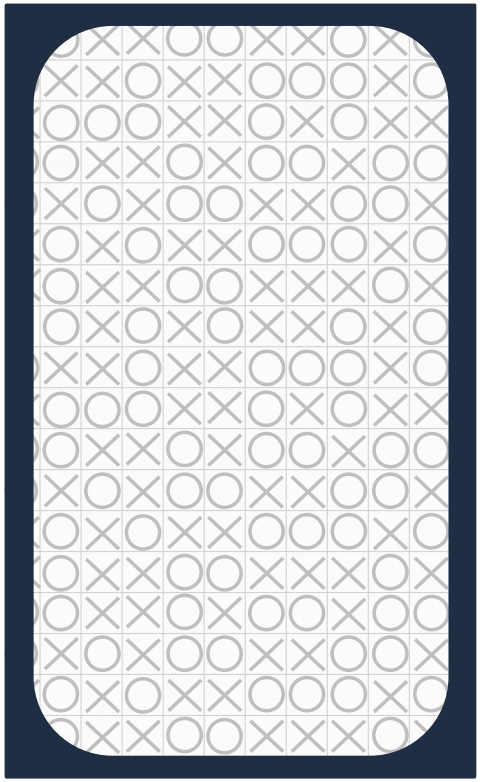
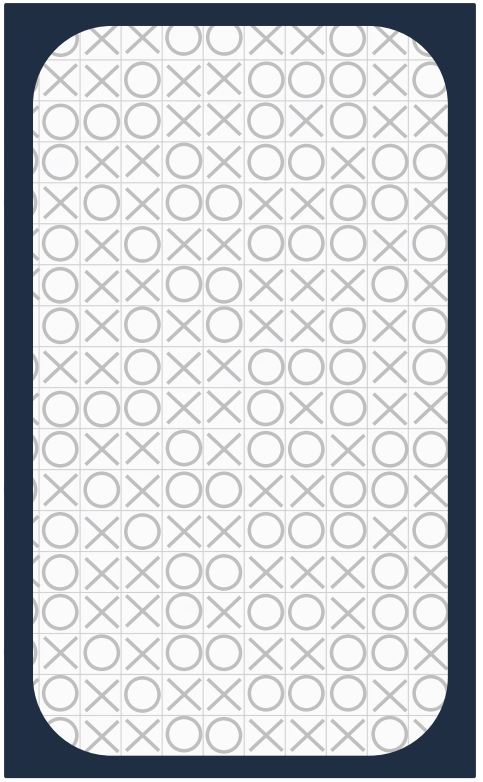
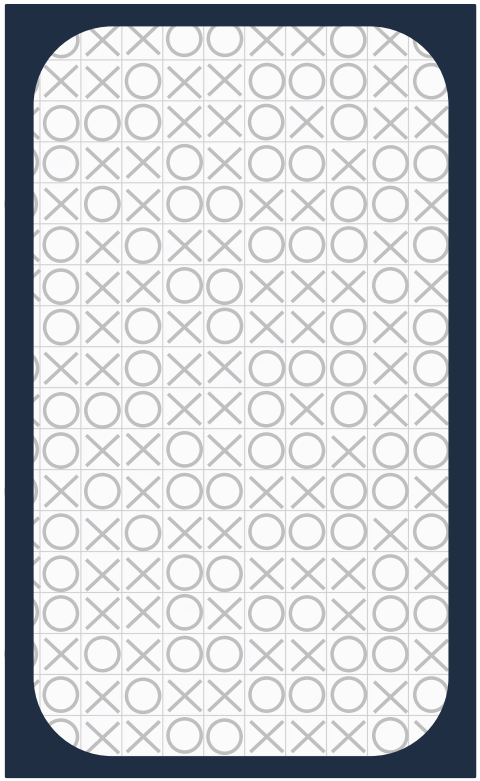
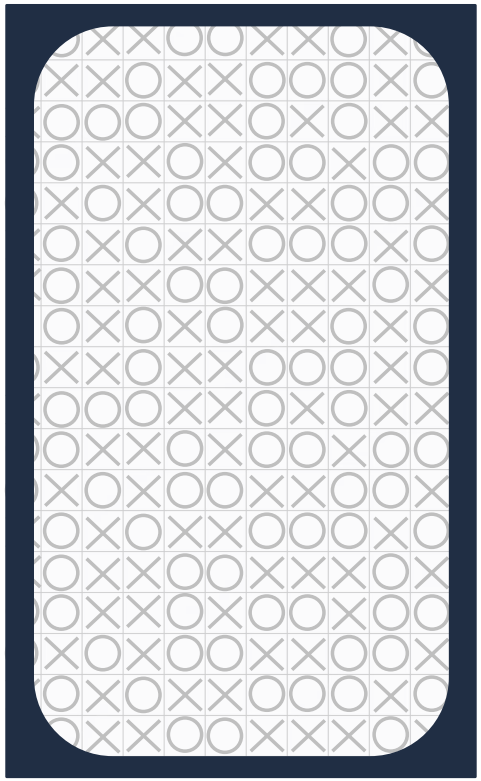
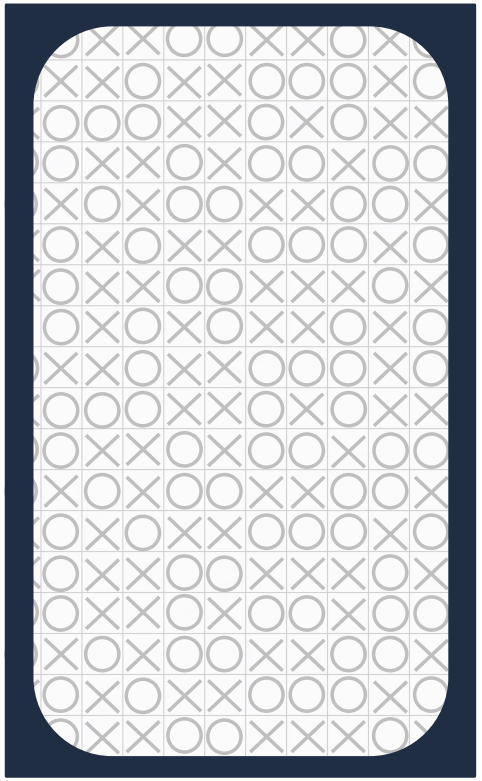
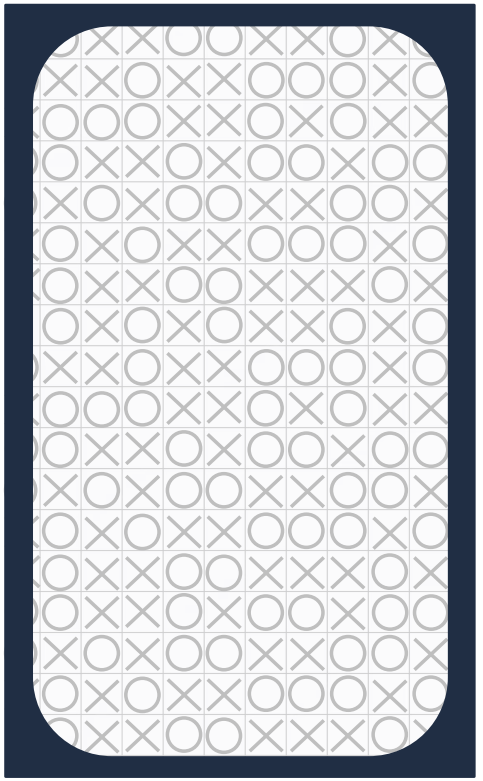


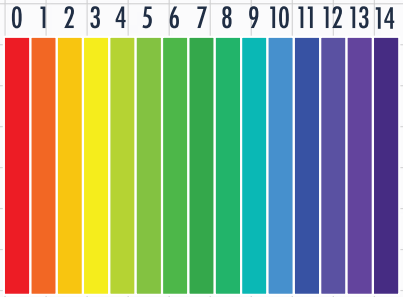
OXIDACE

⇒ ztráta elektronů



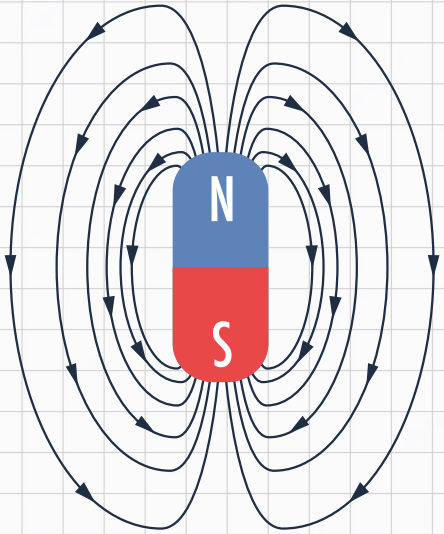









pH



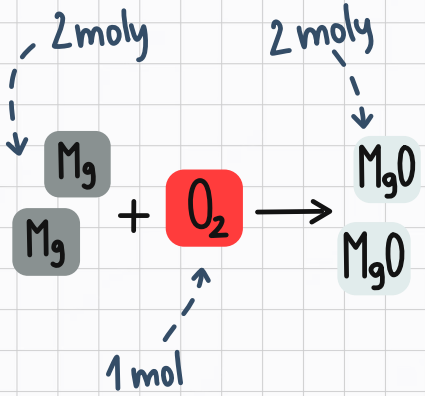
pH škála

POLE



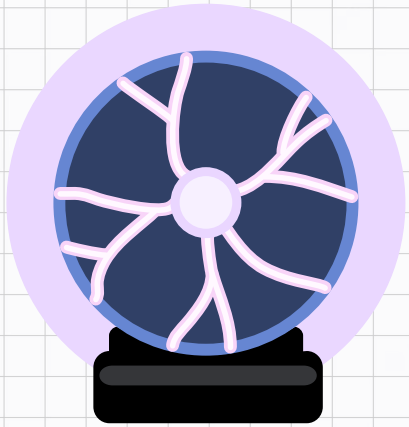
magnetické pole

MOL

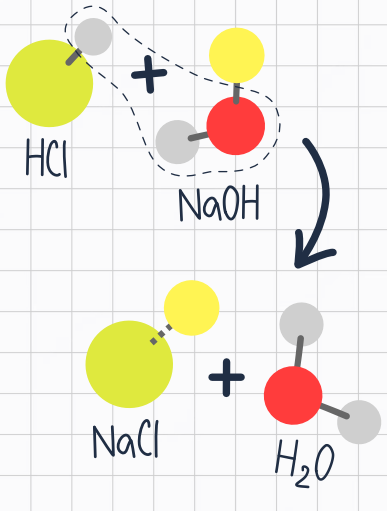


hoření hořčíku

PLAZMA

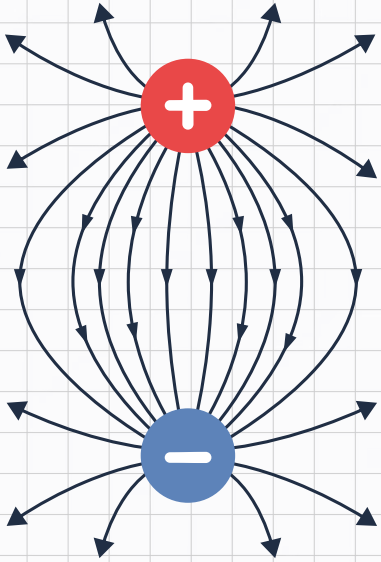


pH



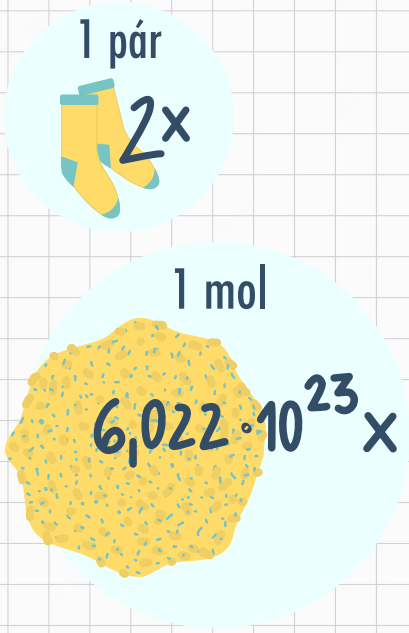
neutralizace

POLE



elektrické pole

MOL



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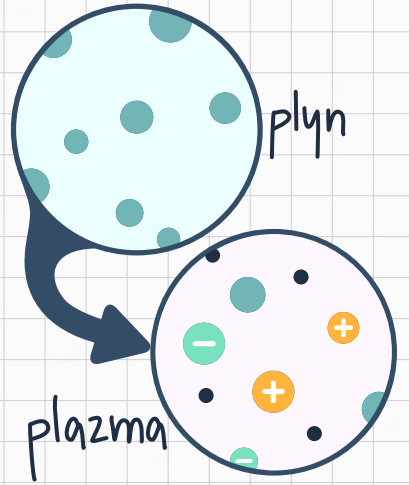
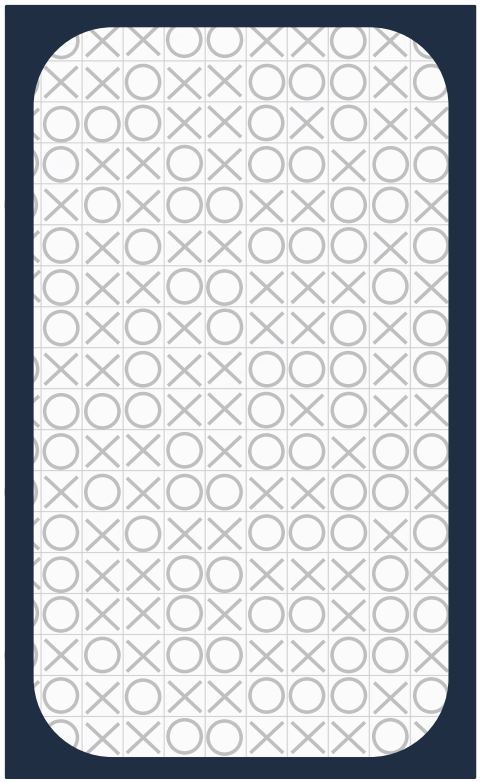
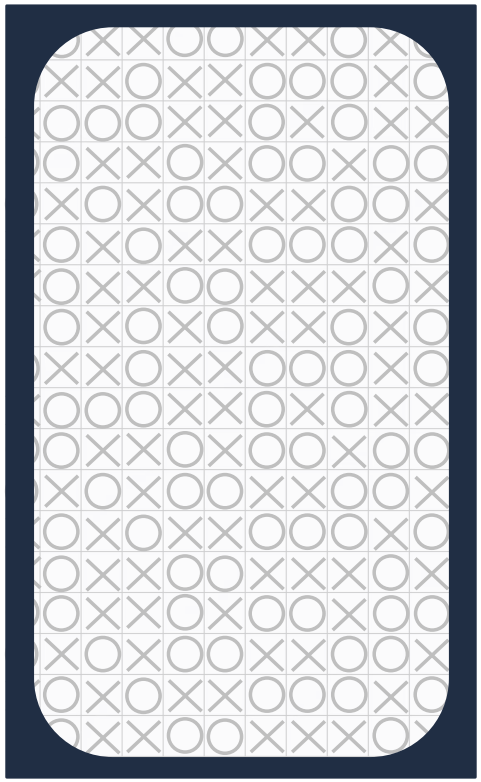
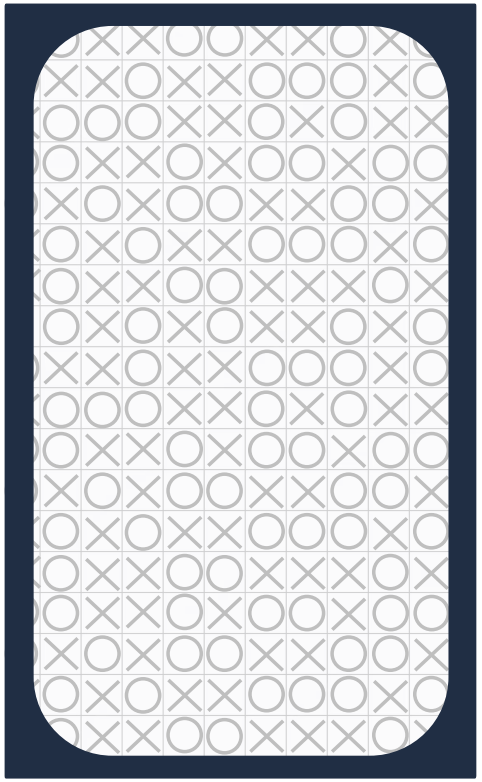
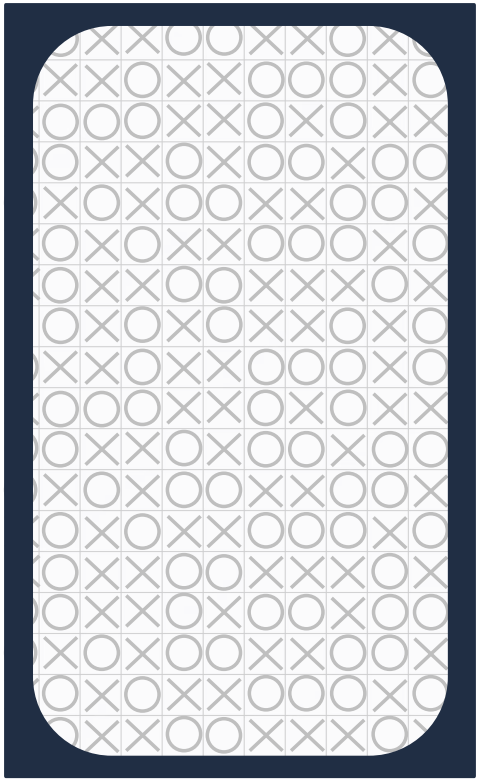
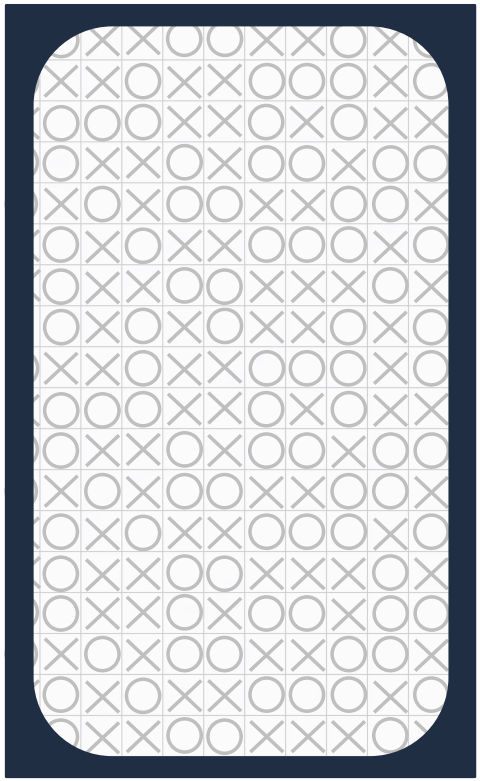
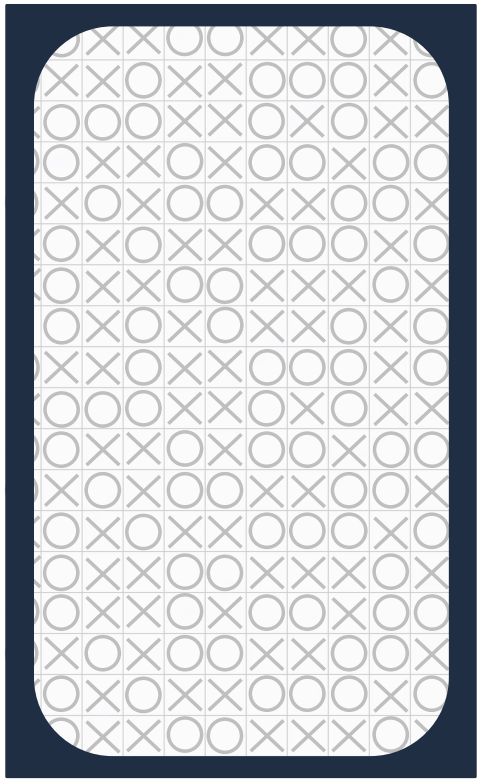
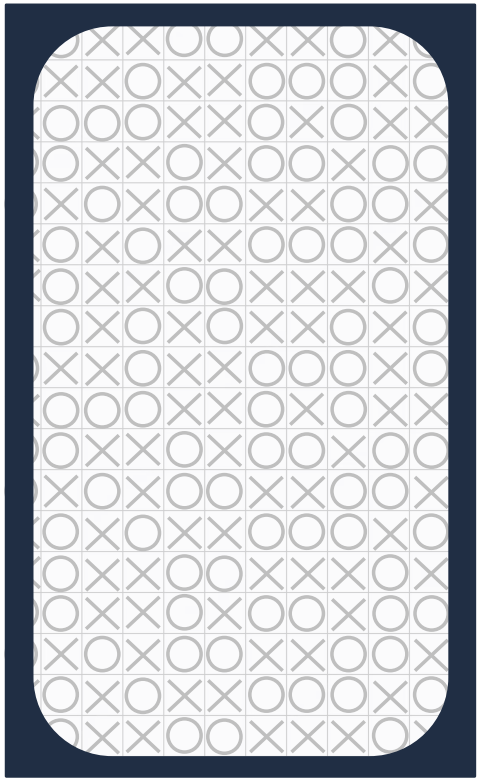
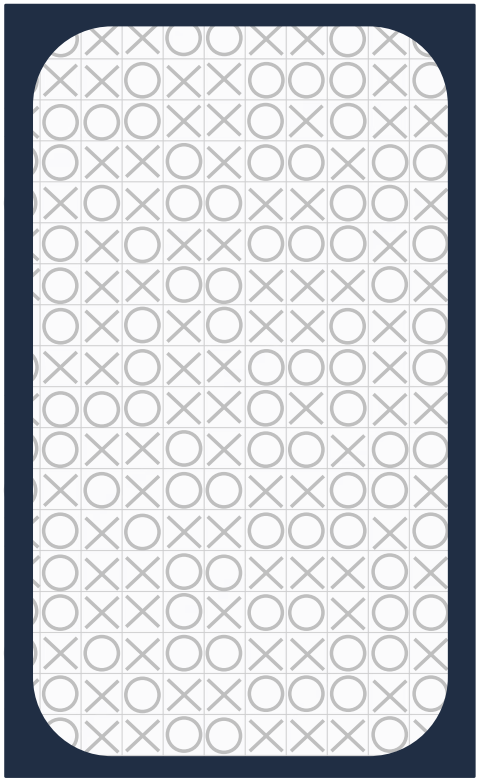


e⁻ ionty molekuly

ČERNÁ DÍRA